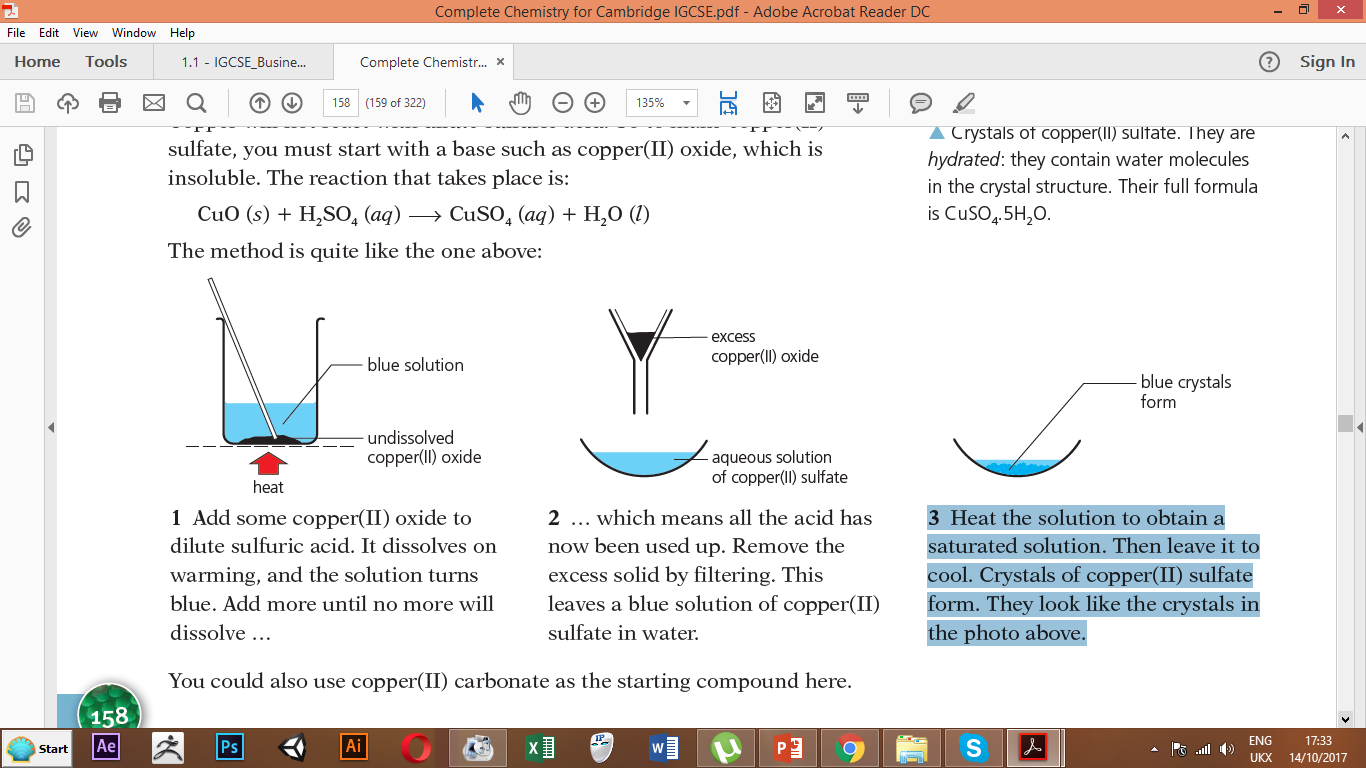
1. What will you start with, to make the salt zinc chloride?  
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2. You would not make lead salts by reacting lead with acids.  
   **a** Why not? **b** Suggest a way to make lead nitrate.  
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3. Look at step 2 on the right. The zinc was in excess.   
   What does that mean? (Check the glossary?)  
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**2** … which means all the acid has now been used up. Remove the excess solid by filtering. This leaves a blue solution of copper(II) sulfate in water.

1. What is the purpose of a titration?  
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2. For carrying out a titration, a burette and pipette are used rather than measuring cylinders. Why?  
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3. You are asked to make the salt ammonium nitrate. Which reactants will you use?  
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